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Monitoring *in situ* the colloidal synthesis of AuRh/TiO₂ selective-hydrogenation nanocatalysts

Zere Konuspayeva,^a Gilles Berhault,^a Pavel Afanasiev,^a Thanh-Son Nguyen,^a Suzanne Giorgio^b and Laurent Piccolo ^{*a}

AuRh nanoparticles (NPs) of various compositions and sizes in the 2–4 nm range were synthesized using a colloidal approach and were characterized at each preparation step by dynamic light scattering (DLS), ultraviolet-visible (UV-vis) spectroscopy, and liquid-phase transmission electron microscopy (liquid TEM). The AuRh colloids appear relatively instable, leading to their gradual coalescence. After fast immobilization of the metallic nanoparticles on rutile TiO₂ nanorods, the materials were investigated by high-resolution transmission electron microscopy (HRTEM) and low-temperature CO adsorption monitored by Fourier transform infrared (FTIR) spectroscopy. The inherent lack of miscibility between Au and Rh leads to partial segregation inside the NPs, which is further exalted after a reducing thermal treatment applied for PVA removal. The catalytic properties in the liquid-phase selective hydrogenation of cinnamaldehyde to hydrocinnamaldehyde are strongly influenced by these nanostructural modifications. While in as-prepared samples the intermixing between Au and Rh phases promotes the catalytic performances for Rh-rich AuRh catalysts through Au-induced stabilization of Rh in its metallic form, segregation into Janus particles after reduction decreases the catalytic activity.

Introduction

Bimetallic nanoparticles are of prominent importance in heterogeneous catalysis since the choices of metals and compositions allow one to tune the catalyst selectivity and improve its activity and stability.¹ Among the noble bimetallic catalysts, AuPd has been the most extensively studied system due both to the miscibility of Au and Pd for all compositions, ensuring formation of a homogeneous alloy structure, and to strong synergetic effects observed in numerous catalytic applications.² On the opposite, the AuRh system has only been rarely investigated. Indeed, Au and Rh show large differences in cohesion and surface energies and a large lattice mismatch resulting in a broad miscibility gap in the bulk AuRh phase diagram.³ However, some attempts have been made to prepare AuRh NPs with controlled composition and structure. Extensive work on this system was performed by Kiss, Óvári and co-workers.^{4–7} Using either chemical or physical preparation techniques, they showed that Au tends to cover the Rh surface. Using scanning transmission electron microscopy (STEM), Chantry *et al.*^{8,9} highlighted the tendency of Rh either to deposit in an uneven way on Au nanorods and/or to form mixed alloy layers at the topmost surface depending on kinetic issues

during the formation of such nanostructures. Similarly, Sneed *et al.*¹⁰ observed the formation of island overgrowth of Rh onto Au nanocubes or octahedra. Using microwave irradiation, Garcia *et al.*¹¹ synthesized AuRh NPs stabilized by PVP (polyvinylpyrrolidone) and deposited on amorphous silica, which exhibited a slight enhancement of activity in cyclohexene hydrogenation as compared to Rh NPs. Essinger-Hileman *et al.*¹² synthesized AuRh solid-solution alloy NPs, though for very low amounts of metal precursors. Recently, Shubin *et al.* successfully prepared metastable AuRh solid solutions from the thermal decomposition of home-made bimetallic precursors.¹³

In spite of the high hydrogenation performance of Rh and the promoting capability of Au,³ selective hydrogenation applications have been rarely envisaged up to now for AuRh bimetallic systems, limiting our knowledge of structure/performance relationships. The selective hydrogenation of cinnamaldehyde (CALD)^{14–16} to hydrocinnamaldehyde (HCALD) is of interest for flavor and fragrance industries.¹⁷ Moreover, the reaction is structure-sensitive, as its performance depends on the type of exposed crystallographic planes,^{18–21} the nature and size of the metallic NPs,^{17,19,22–26} and the nature of the support.^{22,23} Some studies have shown the interest of using TiO₂ as a support for modifying the activity and the selectivity of Au NPs.²⁷ Gold-based bimetallic NPs considered up to now for this reaction are AuPd,^{28,29} AuCu,³⁰ and AuNi³¹ systems.

We have previously reported on the colloidal preparation and characterization of AuRh NPs immobilized on rutile titania nanorods.^{3,32} The model catalysts were evaluated for the

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hydrogenation of tetralin in the presence of H_2S^3 and the hydrodeoxygenation (HDO) of guaiacol,³³ both reactions being performed in the gas phase. In both cases, the bimetallic combination proved to behave differently from the Au and Rh counterparts, and the catalytic performance was improved by Au–Rh nanoalloying. It was also shown that calcination–reduction and direct reduction treatments are able to remove the PVA surfactant for activating the catalysts. However, the former treatment leads to metal phase separation into monometallic particles, while the direct reduction leads to phase segregation inside the bimetallic particles.³ In the latter case, an Au/Rh/TiO₂ stacked structure was evidenced, and further STEM characterization and modeling allowed us to explain this chemical arrangement by the lower surface and cohesion energies of Au with respect to Rh, and the preferential binding of Rh with the titania support.^{34–36}

Here, we report on the nucleation and growth of AuRh NPs directly characterized in solution, the surface characteristics of the corresponding as-prepared or treated titania-supported catalysts, and their performance in liquid-phase selective hydrogenation of cinnamaldehyde. In particular, this study is focused on the relationship between the catalytic properties of the supported NPs and their chemical structure, which is itself influenced by the pretreatment.

Experimental section

Materials synthesis methods

TiO₂ rutile nanorods were prepared using a simplified procedure based on a hydrothermal method reported by Li and Afanasiev.³⁷ 10 g of Degussa P25 TiO₂ (50 m² g⁻¹) and 100 mL of 15 wt% H₂SO₄ solution were mixed in a Teflon reactor and placed in a sealed autoclave kept at 200 °C for 15 days. The obtained solid was washed several times with 0.1 M NH₄NO₃ to remove adsorbed sulfate, then washed with distilled water, dried at 100 °C overnight, and calcined at 350 °C in air for 2 h.

AuRh NPs, along with their monometallic counterparts, were prepared using a colloidal (co)reduction route. The metal precursors were HAuCl₄·3H₂O (Strem Chemicals, 99.9%, 49 wt% Au) and RhCl₃·nH₂O (Sigma-Aldrich, 99.9%, 38–40 wt% Rh). In a first step, a 200 mL aqueous solution containing the two metallic precursors was prepared by adding the amounts of precursors necessary for reaching a total metal loading of 3 wt%. Next, a 1 wt% aqueous solution of polyvinyl alcohol (PVA, $M_w = 10\,000$) was added to the preceding solution while maintaining a mass ratio $m_{\text{PVA}}/m_{\text{Au+Rh}}$ of 1.2. A solution of 0.1 M NaBH₄, freshly prepared and kept at 0 °C before use, was then dropped under stirring in the solution of metallic precursors with a molar ratio $n_{\text{NaBH}_4}/n_{\text{Au+Rh}}$ of 5 to ensure a complete reduction of the two metals. Immediately after NaBH₄ addition, the color of the solution changed from yellowish orange to black, indicating the reduction of Au³⁺ and Rh³⁺ ions into Au⁰ and Rh⁰. Stirring was then maintained for 30 min to allow the complete decomposition of the exceeding NaBH₄. The solution was then acidified to pH 3.5 by addition of HCl 0.01 M in order to favor the sol immobilization onto the TiO₂ support. The amount of support necessary for reaching the final metal

loading was then added and stirring was kept for 3 h. Finally, the material was filtrated, washed with hot distilled water (70 °C) several times, and dried at 100 °C overnight. In some cases, a reductive post-treatment was performed by heating the samples to 350 °C (10 °C min⁻¹, 3 h plateau) in hydrogen flow (5 mL min⁻¹, 1 atm). Table 1 reports the series of synthesized catalysts.

Materials characterization methods

The metal loadings were determined by inductively coupled plasma optical emission spectroscopy (ICP-OES, Horiba Jobin Yvon Activa). Before analysis, the samples were treated with a mixture of H₂SO₄, *aqua regia*, and HF at 250–300 °C. Dynamic light scattering (DLS) experiments were performed on a Wyatt DynaPro NanoStar equipment using a 658 nm laser wavelength and a 90° scattering angle for detection. Au and Rh colloidal solutions were analyzed every 20 min with a 5 s acquisition time while the evolution of AuRh NPs in solution was determined by recording data every 30 min with the same acquisition time. The Dynamics software was used to acquire the data. UV-vis spectra of the PVA-stabilized Au, Rh, and AuRh colloidal suspensions were recorded on a Lambda 35 (Perkin Elmer) spectrometer in the 190–1100 nm wavelength region at 1 nm min⁻¹ acquisition rate. Transmission electron microscopy (TEM) was performed using an aberration-corrected FEI Titan ETEM G2 operated at 300 kV (CLYM). For TEM sample preparation, the catalysts were crushed in ethanol and the solution was ultrasonically stirred before dropping it on a holey carbon-covered copper TEM grid, followed by drying. For liquid-phase TEM experiments, an aqueous solution containing the Au and Rh precursors and the PVA surfactant was encapsulated between two graphene oxide layers on a carbon film deposited onto a copper grid before being observed with a Jeol 3010 TEM operating at 300 kV.³⁸ During the observation at a constant magnification of 10⁵, the current density was stabilized at about 150 pA cm⁻² at the screen level. During the reduction in the electron beam, the nucleation and growth of AuRh NPs was followed in real time in a selected area for a few minutes. In this work, the selected images were recorded before the solvent evaporation, so that the metal concentration in the observed area was constant. Transmission infrared spectroscopy of adsorbed CO (CO-FTIR) was carried out using a Thermo Nicolet 6700 spectrometer (DTGS detector) combined with a custom-made glass cell³

Table 1 Characteristics of the supported catalysts

Catalyst	Au : Rh atomic ratio ^a		Metal loading ^a (wt%)	Particle size ^b (nm)
	Targeted	Measured		
Rh/TiO ₂	0 : 100	—	0.6	2.4 ± 0.6
Au ₁₃ Rh ₈₇ /TiO ₂	10 : 90	13 : 87	2.2	2.4 ± 0.9
Au ₃₇ Rh ₆₃ /TiO ₂	30 : 70	37 : 63	2.4	3.2 ± 1.2
Au ₅₇ Rh ₄₃ /TiO ₂	50 : 50	57 : 43	2.2	4.0 ± 1.3
Au ₆₃ Rh ₃₇ /TiO ₂	60 : 40	63 : 37	2.5	3.3 ± 1.0
Au/TiO ₂	100 : 0	—	1.5	2.8 ± 0.8

^a Determined by ICP-OES. ^b Determined by TEM.

which allows for *in situ* treatment (gas flow, 1 atm, RT-500 °C) of the samples. CO (Air Liquide, 99.997% purity) was dosed using a stainless steel buffer and electro-valves. The samples were prepared in the form of self-supported pellets (13 mm diameter, *ca.* 500 mm thickness, *ca.* 20 mg weight), which were pretreated at 350 °C (5 °C min⁻¹, 1 h plateau) in H₂ flow when needed, and outgassed at 350 °C for 1 h under vacuum. Cumulative dosing of CO (1 μmol per dose) and subsequent spectrum acquisition were performed until saturation coverage. The spectra were recorded with 128 scans and 4 cm⁻¹ resolution in the 4000–900 cm⁻¹ range.

The hydrogenation of cinnamaldehyde in liquid phase was performed in a stirred autoclave. In a typical experiment, 2 mmol of cinnamaldehyde (99%) were dissolved into 100 mL of 2-propanol and purged twice with N₂. The catalysts were then introduced into the autoclave before a next N₂ purge. The reaction was performed at 50 °C under 10 bar H₂ pressure at a stirring speed of 800 rpm, for 7 h. A small amount (0.3 mL) of the reacting mixture was withdrawn every 30 min and analyzed by gas chromatography (Varian 3380).

Results and discussion

Characterization of colloidal suspensions

Dynamic light scattering. The first goal of our study was to determine the particle sizes and their stability for colloidal suspensions of Au, Rh and AuRh (50 : 50 mol% target composition) particles stabilized by polyvinyl alcohol (PVA). DLS allows evaluating *in situ* the hydrodynamic diameter³⁹ of as-formed colloidal particles and their size evolution as a function of time. As shown in Fig. 1, as soon as the NaBH₄ reducing agent is added ($t = 0$), colloids are immediately obtained due to a fast nucleation process. The Au colloids exhibit an initial dynamic size of 6.5 nm increasing only slightly with time (up to 7.0 nm after 13 h in solution). They are therefore quite stable, suggesting that gold NP nucleation is fast, with complete

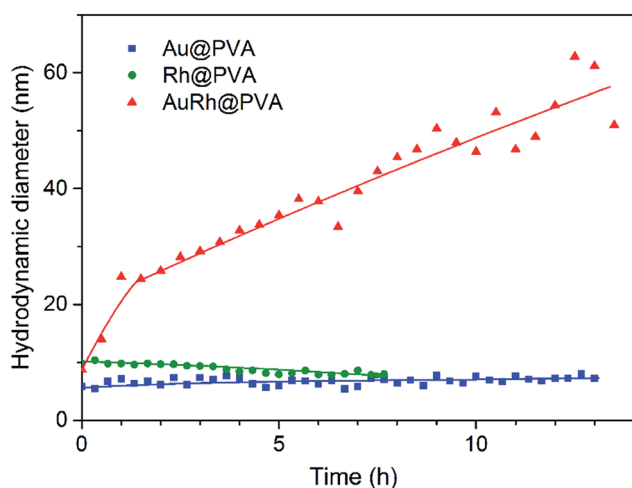


Fig. 1 Evolution of the hydrodynamic diameter (determined by DLS) of PVA-stabilized Au, Rh, and AuRh colloids during their synthesis in solution.

consumption of the precursor and no further Ostwald ripening. The size evolution of Rh colloids somewhat differs from that of Au colloids. A slow but continuous decrease in size is observed. The initial size of about 10 nm gradually decreases down to 8 nm after 5 h in solution. This suggests a consumption of Rh NPs by corrosive oxidation. Indeed, the O₂/Cl⁻ mixture (chloride coming from the Au and Rh precursors), known to act as an oxidative corroding agent,⁴⁰ is present under our experimental conditions. This likely causes the depletion of Rh revealed by elemental analysis of the supported catalysts (see below). As compared to the Rh situation, the absence of any corrosive oxidation for the Au NPs indicates a stronger stabilizing effect of PVA and/or a lower reactivity in the Au case. It should be noticed that after 48 h in solution, no further size evolution of the monometallic colloids can be observed. The AuRh NPs present a rapid growth in size, suggesting a fast aggregation process. Indeed, they exhibit an initial dynamic size of 8–10 nm, while after 10 h in solution, the average size reaches up to 50 nm. The stability of the bimetallic AuRh colloids is thus strikingly lower than that of the monometallic sols, which can be related to the lack of miscibility between Au and Rh and/or an incomplete protecting effect of polyvinyl alcohol (PVA) ligands. Therefore, in the next step of sol immobilization on the TiO₂ support, the fresh AuRh sols were used after only 1 h ageing in order to maintain their final dynamic size below 10 nm.

***In situ* liquid transmission electron microscopy.** The colloidal growth of AuRh NPs was also monitored *in situ* inside a TEM apparatus enabling their visualization in the liquid phase. Fig. 2a displays the evolution of AuRh NPs in solution during the reduction induced by the electron beam. The nucleation step is fast and the NPs are already visible after 20 seconds of irradiation. The average particle size increases up to 4 nm (Fig. 2b), consistently with the value determined from *ex situ* TEM (see below). A previous STEM investigation of unsupported AuRh@PVA colloids has shown that the NPs exhibit roundish shapes and several structural types (single-crystalline, multi-twinned, *etc.*)³⁴ The density of NPs as a function of the irradiation time is reported in Fig. 2c. The number of AuRh NPs first increases rapidly and reaches a plateau after 25 seconds, then it decreases due to the frequent coalescence of NPs in solution, as also directly visible in the image series of Fig. 2a and 3. The growth mechanism corresponds to the direct adsorption of monomers³⁸ during 25 s, followed by coalescence and/or Ostwald ripening, which probably result from a partial capping of the metal by PVA, as also observed *in situ* for Pt NPs.⁴¹ These processes allow stabilizing the AuRh NPs by reducing their surface energy through their increase in size.⁴² This result agrees with the DLS measurements, showing the relative instability of AuRh bimetallic sols.

UV-vis spectroscopy. Au, Rh and AuRh sols were also characterized by UV-vis spectroscopy (Fig. 4). The Au NPs exhibit a clear plasmon band at 550 nm characteristic of spherical particles.⁴³ In contrast, the Rh NPs only show a weak but discernible plasmon band in the UV region, with a maximum around 230 nm and a tail extending from 300 nm to 900 nm.⁴⁴ The addition of Rh to Au leads to the complete disappearance of the Au plasmon band at 550 nm, while the Rh plasmon can still

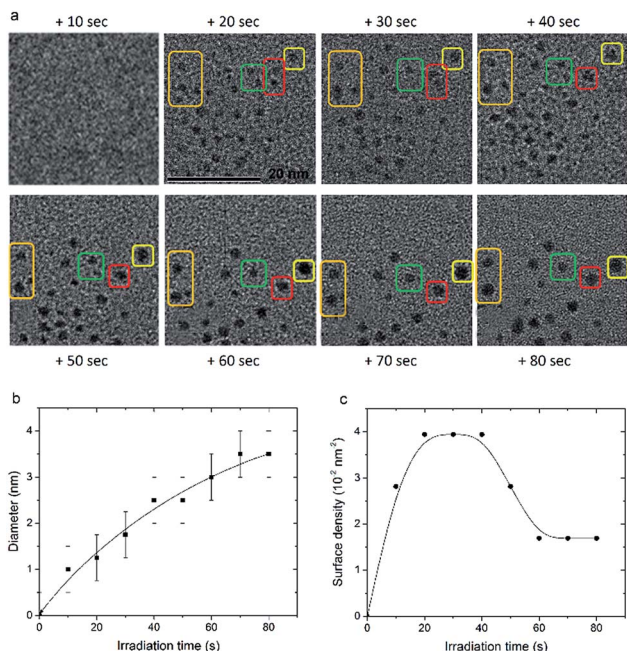


Fig. 2 (a) *In situ* liquid TEM images showing the nucleation and growth of AuRh nanoalloys in solution during reduction of the precursors by the electron beam. The colored rectangles allow following individual particles. The actual sample size for all images is 30×30 nm. The bottom graphs show the corresponding NP diameter (b) and surface density (c) in the solution as a function of electron irradiation time.

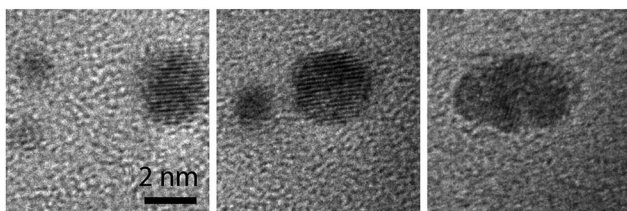


Fig. 3 Successive liquid TEM images showing the gradual coalescence of AuRh NPs in solution (from left to right). The images were recorded with a time interval of 10 s.

be detected around 230 nm. This supports the formation of bimetallic AuRh NPs. Moreover, the detection of the Rh plasmon suggests the formation of Au@Rh core-shell NPs, as claimed by Toshima.⁴⁵ However, post-synthesis STEM observations have shown that several shapes and chemical configurations of AuRh NPs co-exist.³⁴

Characterization of supported catalysts

Elemental analysis. After formation of the metallic NPs in solution, the sols were deposited onto rutile TiO_2 nanorods through acid-assisted immobilization using a 0.01 M HCl solution, as previously reported.³ Table 1 gathers the metal loadings, the Au : Rh atomic ratios (determined by ICP-OES), and the average NP sizes (TEM) for the Au, Rh, and AuRh catalysts. As previously mentioned, a slight deficiency in Rh is systematically observed with respect to the target loadings. This

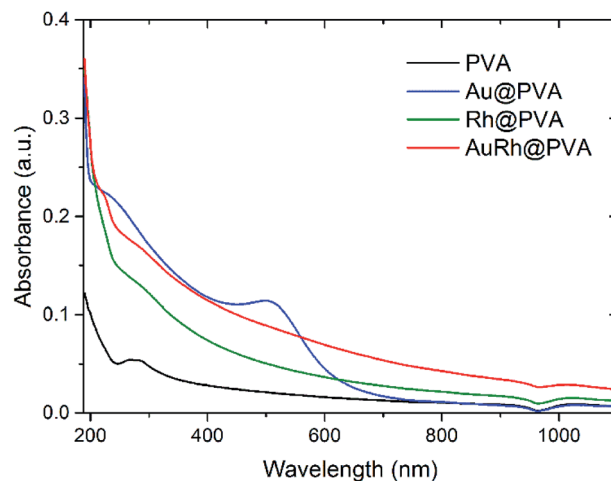


Fig. 4 UV-vis spectra of the Au, Rh, and AuRh sols stabilized by PVA. For comparison purpose, the UV-vis spectrum of PVA alone is also reported (at the same concentration as that used for sols of metallic NPs).

could be related to the partial dissolution (leaching) of Rh incorporated in the AuRh NPs through the corrosive oxidation already suggested by DLS measurements.

Transmission electron microscopy. As shown in the representative examples of Fig. 5, the catalysts contain small-sized metallic NPs distributed over rutile TiO_2 nanorods. In the case of $\text{Au}_{63}\text{Rh}_{37}/\text{TiO}_2$, bimetallic NPs, inhomogeneous in shape and often containing single-phase (Au-rich or Rh-rich) domains, coexist with smaller monometallic Rh NPs.³ The supported particles appear more faceted than the unsupported ones, and truncated at the interface with the support (due to metal wetting

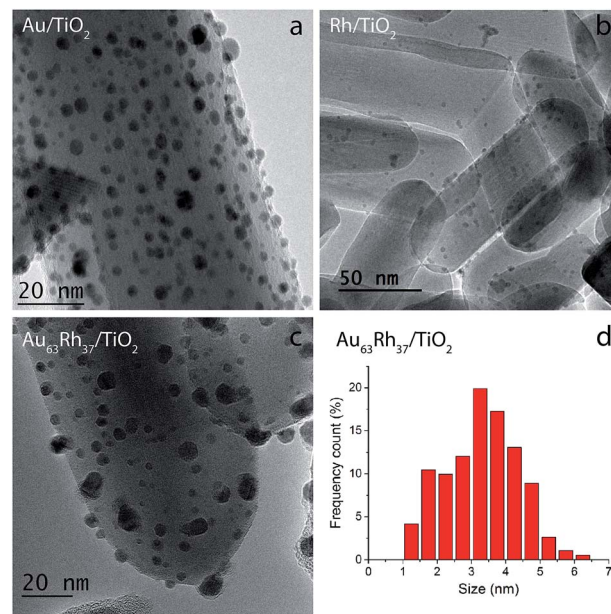


Fig. 5 Representative TEM images of Au/ TiO_2 (a), Rh/ TiO_2 (b) and $\text{Au}_{63}\text{Rh}_{37}/\text{TiO}_2$ (c) as-prepared samples. (d) AuRh particle size distribution for the sample depicted in (c).

over the support). One should note the importance of the nature of the acid used for the sol immobilization.^{3,46} Using HCl as in the present case leads to well-dispersed bimetallic NPs, whereas H₂SO₄ was found to generate a lower proportion of bimetallic NPs and the presence of sulfur impurities.³² The influence of the relative Au/Rh molar ratio on the metal particle size in the AuRh/TiO₂ catalysts was also determined by TEM (Table 1). A gradual increase in the particle size with the Au concentration is evidenced, going from 2.4 nm for Rh and Au₁₃Rh₈₇, to 3.2 nm for Au₃₇Rh₆₃, and 4.0 nm for Au₅₇Rh₄₃. Above 60 at% Au, the average particle size somewhat decreases to 3.3 nm for Au₆₃Rh₃₇, and 2.8 nm for pure Au. As previously reported,^{3,34} after treatment in H₂ flow at 350 °C, the bimetallic NPs keep roughly the same size (3.7 ± 1.3 nm instead of 3.3 ± 1.0 nm for the Au₆₃Rh₃₇ sample) but adopt a Janus configuration with the Rh side contacting the titania support. The (fcc) Au side has a truncated-octahedral morphology, and thus exhibits (111) and (100) facets.³⁴

CO-FTIR spectroscopy. Au, Rh, and AuRh nanocatalysts supported on TiO₂, as well as the bare support, were characterized by low-temperature CO adsorption monitored by FTIR

spectroscopy. As shown in Fig. 6, the samples were either analyzed as-prepared (“fresh”) or after *in situ* reduction in H₂ flow at 350 °C. In the case of the fresh TiO₂ support, a main contribution at 2181 cm⁻¹ due to the chemisorption of CO on β'' Ti⁴⁺ Lewis acid sites is present, while the 2158 cm⁻¹ band is related to CO physisorption on Ti-OH sites, and the 2146 and 2137 cm⁻¹ bands correspond to CO physisorption on TiO₂.⁴⁷⁻⁴⁹ The freshly prepared Au/TiO₂ sample presents similar bands, although the proportion of the CO-Ti⁴⁺ chemisorption band is smaller than for bare titania. Finally, fresh Rh/TiO₂ and AuRh/TiO₂ only exhibit CO-TiO₂ physisorption bands. After the reductive treatment, all the bands related to CO adsorption on titania are present, with the proportion of CO-Ti⁴⁺ increasing in the order: TiO₂ < Au/TiO₂ < AuRh/TiO₂ < Rh/TiO₂. Moreover, the bands corresponding to CO chemisorption on Au⁰ (2102-2105 cm⁻¹), Rh⁰ (2069 cm⁻¹), and both Au⁰ (ca. 2100 cm⁻¹) and Rh⁰ (2067 cm⁻¹)⁴⁸⁻⁵⁵ are observed for the Au/TiO₂, Rh/TiO₂, and AuRh/TiO₂ catalysts, respectively. However, as the CO-Au contribution is very small in the case of AuRh/TiO₂, Rh may form a thin surface layer surrounding the Au region of the Janus particles. This would correspond to an

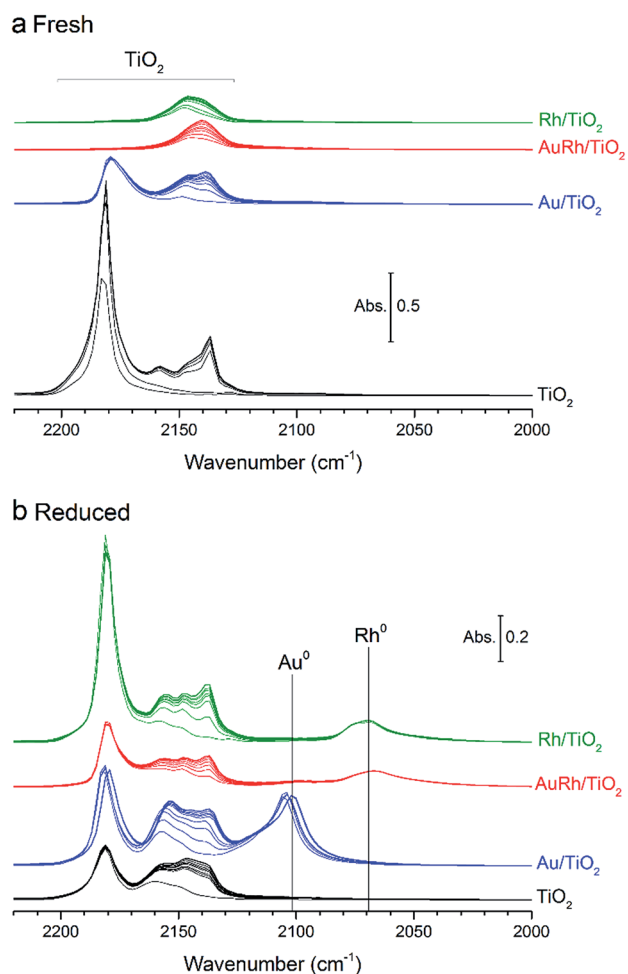


Fig. 6 CO-FTIR spectra of TiO₂, Au/TiO₂, Rh/TiO₂ and Au₆₃Rh₃₇/TiO₂ as-prepared (a) and *in situ* pre-reduced (b) catalysts. The spectra were recorded at increasing CO coverages (2 to 42 μmol CO doses).

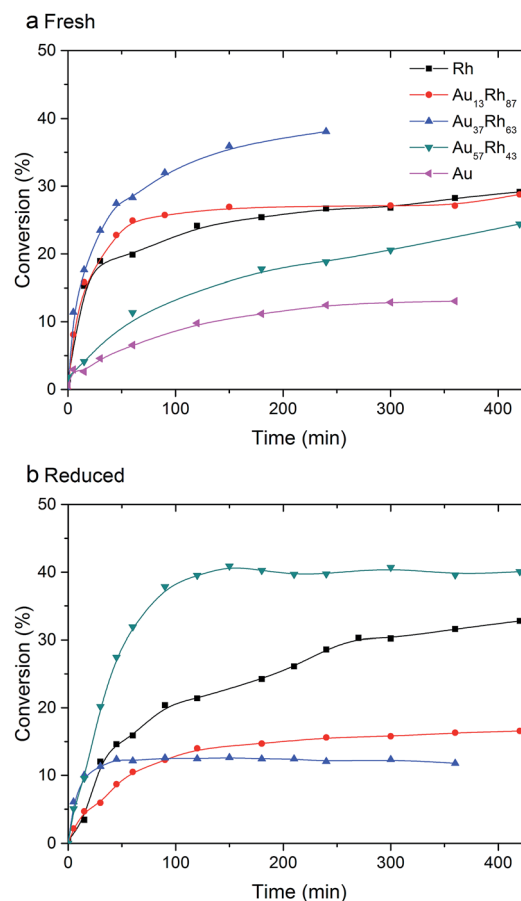


Fig. 7 Evolution of conversions in cinnamaldehyde hydrogenation as a function of the reaction time for fresh (a) and pre-reduced (b) AuRh/TiO₂ catalysts. The reaction was performed at 50 °C under 10 bar H₂ at a stirring speed of 800 rpm.

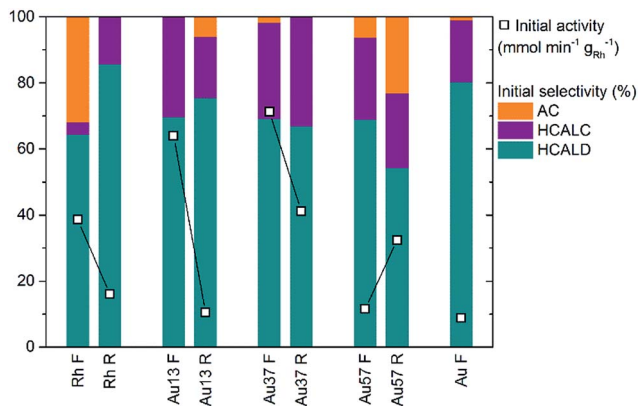


Fig. 8 Initial activities (square dots) and selectivities (bars) in cinnamaldehyde hydrogenation for fresh (F) and pre-reduced (R) AuRh/TiO₂ catalysts. AC, HCALC and HCALD denote hydrocinnamaldehyde diisopropyl acetal, hydrocinnamyl alcohol and hydrocinnamaldehyde, respectively. The initial activity is expressed in mmol min⁻¹ g_{Rh}⁻¹ except for Au/TiO₂ (mmol min⁻¹ g_{Au}⁻¹). Reaction conditions same as Fig. 7.

off-centered Au core inside a Rh shell, as already observed for, e.g., the Ag-Cu system.⁵⁶

Catalytic performances

Cinnamaldehyde (CALD) hydrogenation was selected as a model reaction to evaluate the catalytic properties of AuRh/TiO₂ systems in comparison to their monometallic counterparts, both as-prepared and after the reductive treatment. The main products detected under our experimental conditions were hydrocinnamaldehyde (HCALD, desired saturated aldehyde) and 3-phenyl-1-propanol, *i.e.* hydrocinnamyl alcohol (HCALC, undesired saturated alcohol). Cinnamyl alcohol (CALC) was not formed under our experimental conditions. The only (undesired) by-product detected in a non-negligible amount was an acetal (AC) compound, namely

hydrocinnamaldehyde diisopropyl acetal, which is formed by reaction of HCALD with the 2-propanol solvent (see below).²⁴ The catalysis results are reported in Fig. 7 (conversions vs. time), Fig. 8 (initial activities and selectivities) and Table 2 (numeric data).

Strikingly, the fresh Rh-rich catalysts (Rh molar concentrations of 100%, 87% and 63%) appear much more active than their pre-reduced counterparts. This indicates that the presence of PVA at the surface of the particles does not inhibit the reaction, or that the surfactant is removed through *in situ* washing by the solvent. The initial activities of the fresh catalysts rank in the order: Au₃₇Rh₆₃ > Au₁₃Rh₈₇ > Rh ≫ Au₅₇Rh₄₃ > Au. The most active catalyst, Au₃₇Rh₆₃/TiO₂, presents an initial activity of 71 mmol min⁻¹ g_{Rh}⁻¹ (the activities have been calculated per Rh weight, except for pure Au, since Rh is intrinsically much more active than Au) and a conversion of 37% after 3 h of reaction. The corresponding initial turnover frequency is estimated to 0.57 CALD molecule converted per surface Rh atom per second, which is close to the values (0.50–0.61 s⁻¹) reported for Rh/C catalysts under similar conditions.¹⁷ After reduction, while all the Rh-rich catalysts are partially deactivated, Au₃₇Rh₆₃/TiO₂ remains the catalyst with the highest initial activity (41 mmol min⁻¹ g_{Rh}⁻¹), but rapidly deactivates. Au₅₇Rh₄₃/TiO₂ is the only catalyst showing an activation upon reduction, the initial activity changing from 12 to 32 mmol min⁻¹ g_{Rh}⁻¹, and it exhibits the highest final conversion (40%).

For the fresh samples, the initial selectivities to HCALD are similar for all the Rh-containing catalysts (64–70%). Only Au/TiO₂ presents a higher selectivity (80%) but its activity is quite low. HCALC forms with a very low selectivity on Rh/TiO₂ (4%), while higher selectivities are measured for Au-containing catalysts (19–30%). Interestingly, unlike for the Rh/TiO₂ catalyst, the acetal by-product is initially not or almost not formed on AuRh/TiO₂ catalysts. Au₁₃Rh₈₇/TiO₂ presents the highest HCALC selectivity at 20% conversion (69%, Table 2) and no acetal formation. After reduction, the initial selectivities to HCALD

Table 2 Catalytic performances^a

Catalyst state	Composition (at%)	Initial activity (mmol min ⁻¹ g _{Rh} ⁻¹)	Initial selectivity (%)			Conversion after 3 h (%)	Selectivity at 20% conversion (%)		
			HCALD	HCALC	AC		HCALD	HCALC	AC
Fresh	Rh	39	64	4	32	25	61	10	29
	Au ₁₃ Rh ₈₇	64	70	30	0	27	69	31	0
	Au ₃₇ Rh ₆₃	71	69	29	2	37	64	26	10
	Au ₅₇ Rh ₄₃	12	69	25	6	18	52	18	30
	Au	9	80	19	1	11	48	27	25
Reduced	Rh	16	86	14	0	24	69	15	16
	Au ₁₃ Rh ₈₇	10	76	18	6	15	61	18	21
	Au ₃₇ Rh ₆₃	41	67	33	0	12	69	31	0
	Au ₅₇ Rh ₄₃	32	54	23	23	40	35	14	51

^a The initial activity is expressed in mmol min⁻¹ g_{Rh}⁻¹ except for Au/TiO₂ (mmol min⁻¹ g_{Au}⁻¹), the initial selectivity is defined as the ratio between the rate of formation of a given product and the rate of consumption of the reactant at *t* = 0, and the selectivity is defined as the ratio between the concentration of a given product and the concentration of all products. For Au/TiO₂, the results are given only without post-treatment, since otherwise this catalyst presented negligible conversion values. HCALD, HCALC and AC denote hydrocinnamaldehyde, hydrocinnamyl alcohol, and hydrocinnamaldehyde diisopropyl acetal, respectively. The reaction was performed at 50 °C under 10 bar H₂ at a stirring speed of 800 rpm.

decrease with increasing Au loading. Noticeably, the Rh/TiO₂ catalyst no longer presents any initial formation of acetal, contrary to its fresh counterpart. In contrast, the Au-rich Au₅₇Rh₄₃/TiO₂ sample exhibits a high initial selectivity to acetal (23%). The best compromise in terms of initial activity and HCALD selectivity is observed for Au₃₇Rh₆₃/TiO₂, but the final CALD conversion is low (12%).

In summary, these results clearly show the superior performances of the Rh-rich fresh AuRh catalysts. Two groups of AuRh systems can be defined: up to ~40 at% Au, the catalysts exhibit behaviors close to that of Rh but with enhanced activity and selectivity to HCALD and decreased acetal formation, while increasing the Au concentration above 50 at% leads to a loss of the beneficial effect of Au addition to Rh. Only for Au₅₇Rh₄₃/TiO₂, the reductive treatment increases the activity, but also favors acetal production.

The formation of by-products in cinnamaldehyde hydrogenation has been extensively reported.^{57–62} They generally result from the reaction of solvents with the hydrocarbon reactant or its hydrogenated, isomerized, or C–O hydrogenolyzed products.^{24,61} Acetals are generally formed when using light alcohols as solvents. However, Rylander and Himelstein have shown that the use of light alcohols is beneficial to the activity.⁶³ Acetal formation is also favored by the Lewis acid character of the catalyst.⁴ In our case, CO-FTIR has shown the presence of Lewis acidity on the catalysts. It may be associated not only to the TiO₂ support but also to Rh^{δ+} sites (*i.e.*, oxidic Rh species), the amount of which was previously shown to decrease in bimetallics due to Au-induced Rh⁰ stabilization.³ Consistently, for the Rh-rich samples, the benefit of a relatively intimate contact between Au and Rh is lost after treatment, due to enhanced metal segregation (Janus NPs). On the opposite, for Au-rich AuRh, the high dilution of Rh with Au in fresh catalysts is consistent with a poor hydrogenation ability, while the formation of Janus-type particles upon treatment would make Rh available for participating in the catalytic process. Finally, since the average metal particle size varies in the range 2.4–4.0 nm (Table 1), some contribution of size and/or morphology effects cannot be excluded, although essentially on the products distribution.¹⁷

Conclusions

The growth of AuRh NPs and the influence of the post-treatment on structural and hydrogenation properties of AuRh/TiO₂ catalysts have been studied using a combination of advanced characterization techniques and catalytic testing. Liquid-phase selective hydrogenation of cinnamaldehyde provides an alternative tool for establishing structure-composition-activity relationships.

DLS and liquid TEM experiments enable an *in situ* monitoring of the nanoparticle ripening and coalescence processes. It was shown that AuRh colloids are much less stable in solution than Au and Rh colloids, which imposes a thorough control of the preparation parameters for obtaining small AuRh NPs (2–4 nm). In fresh Rh-rich bimetallic catalysts, the intermixing between Au and Rh phases favors Au-induced stabilization of Rh in its metallic form, improving both the activity and the

selectivity to hydrocinnamaldehyde with respect to Rh/TiO₂. The reducing treatment at 350 °C, in spite of its ability to remove the PVA surfactant, enhances Au–Rh segregation phenomena and thereby cancels the stabilizing effect of Au, leading to lower catalytic efficiency.

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References

- 1 L. Piccolo, in *Nanoalloys: Synthesis, Structure and Properties*, ed. D. Alloyeau, C. Mottet and C. Ricolleau, Springer London, 2012, pp. 369–404.
- 2 G. J. Hutchings and C. J. Kiely, *Acc. Chem. Res.*, 2013, **46**, 1759–1772.
- 3 Z. Konuspayeva, P. Afanasiev, T.-S. Nguyen, L. D. Felice, F. Morfin, N.-T. Nguyen, J. Nelayah, C. Ricolleau, Z. Y. Li, J. Yuan, G. Berhault and L. Piccolo, *Phys. Chem. Chem. Phys.*, 2015, **17**, 28112–28120.
- 4 J. Raskó, Á. Koós, K. Baán and J. Kiss, *React. Kinet. Catal. Lett.*, 2007, **90**, 187–195.
- 5 L. Óvári, L. Bugyi, Z. Majzik, A. Berkó and J. Kiss, *J. Phys. Chem. C*, 2008, **112**, 18011–18016.
- 6 L. Óvári, A. Berkó, R. Gubó, Á. Rácz and Z. Kónya, *J. Phys. Chem. C*, 2014, **118**, 12340–12352.
- 7 L. Óvári, A. Berkó, G. Vári, R. Gubó, A. P. Farkas and Z. Kónya, *Phys. Chem. Chem. Phys.*, 2016, **18**, 25230–25240.
- 8 R. L. Chantry, W. Siritwacharapiboon, S. L. Horswell, A. J. Logsdail, R. L. Johnston and Z. Y. Li, *J. Phys. Chem. C*, 2012, **116**, 10312–10317.
- 9 R. L. Chantry, I. Atanasov, W. Siritwacharapiboon, B. P. Khanal, E. R. Zubarev, S. L. Horswell, R. L. Johnston and Z. Y. Li, *Nanoscale*, 2013, **5**, 7452–7457.
- 10 B. T. Sneed, C.-H. Kuo, C. N. Brodsky and C.-K. Tsung, *J. Am. Chem. Soc.*, 2012, **134**, 18417–18426.
- 11 S. García, L. Zhang, G. W. Piburn, G. Henkelman and S. M. Humphrey, *ACS Nano*, 2014, **8**, 11512–11521.
- 12 E. R. Essinger-Hileman, D. DeCicco, J. F. Bondi and R. E. Schaak, *J. Mater. Chem.*, 2011, **21**, 11599–11604.
- 13 Y. Shubin, P. Plyusnin, M. Sharafutdinov, E. Makotchenko and S. Korenev, *Nanotechnology*, 2017, **28**, 205302.
- 14 S. Fleischer, S. Zhou, K. Junge and M. Beller, *Angew. Chem., Int. Ed.*, 2013, **52**, 5120–5124.
- 15 G.-H. Wang, X. Deng, D. Gu, K. Chen, H. Tüysüz, B. Spliethoff, H.-J. Bongard, C. Weidenthaler, W. Schmidt and F. Schüth, *Angew. Chem., Int. Ed.*, 2016, **55**, 11101–11105.
- 16 M. Zhao, K. Yuan, Y. Wang, G. Li, J. Guo, L. Gu, W. Hu, H. Zhao and Z. Tang, *Nature*, 2016, **539**, 76–80.
- 17 P. Gallezot and D. Richard, *Catal. Rev.: Sci. Eng.*, 1998, **40**, 81–126.
- 18 F. Delbecq and P. Sautet, *J. Catal.*, 1995, **152**, 217–236.
- 19 P. Claus, *Top. Catal.*, 1998, **5**, 51–62.

- 20 C. J. Kliewer, M. Bieri and G. A. Somorjai, *J. Am. Chem. Soc.*, 2009, **131**, 9958–9966.
- 21 A. Mehri, H. Kochkar, S. Daniele, V. Mendez, A. Ghorbel and G. Berhault, *J. Colloid Interface Sci.*, 2012, **369**, 309–316.
- 22 B. Coq, P. S. Kumbhar, C. Moreau, P. Moreau and M. G. Warawdekar, *J. Mol. Catal. A: Chem.*, 1993, **85**, 215–228.
- 23 M. Englisch, A. Jentys and J. A. Lercher, *J. Catal.*, 1997, **166**, 25–35.
- 24 M. Lashdaf, A. O. I. Krause, M. Lindblad, M. Tiitta and T. Venäläinen, *Appl. Catal., A*, 2003, **241**, 65–75.
- 25 M. E. Grass, R. M. Rioux and G. A. Somorjai, *Catal. Lett.*, 2009, **128**, 1–8.
- 26 X. Yang, A. Wang, X. Wang, T. Zhang, K. Han and J. Li, *J. Phys. Chem. C*, 2009, **113**, 20918–20926.
- 27 P. Claus, A. Brückner, C. Mohr and H. Hofmeister, *J. Am. Chem. Soc.*, 2000, **122**, 11430–11439.
- 28 X. Yang, D. Chen, S. Liao, H. Song, Y. Li, Z. Fu and Y. Su, *J. Catal.*, 2012, **291**, 36–43.
- 29 T. Szumelda, A. Drelinkiewicz, R. Kosydar and J. Gurgul, *Appl. Catal., A*, 2014, **487**, 1–15.
- 30 X. Yuan, J. Zheng, Q. Zhang, S. Li, Y. Yang and J. Gong, *AIChE J.*, 2014, **60**, 3300–3311.
- 31 M. G. Prakash, R. Mahalakshmy, K. R. Krishnamurthy and B. Viswanathan, *Catal. Today*, 2016, **263**, 105–111.
- 32 Z. Konuspayeva, G. Berhault, P. Afanasiev, T.-S. Nguyen, A. Auyezov, M. Burkitbayev and L. Piccolo, in *MRS Proceedings*, MRS Online Proceedings Library, 2014, vol. 1641, pp. mrsf13-1641-aa08-07.
- 33 T.-S. Nguyen, D. Laurenti, P. Afanasiev, Z. Konuspayeva and L. Piccolo, *J. Catal.*, 2016, **344**, 136–140.
- 34 L. Piccolo, Z. Y. Li, I. Demiroglu, F. Moyon, Z. Konuspayeva, G. Berhault, P. Afanasiev, W. Lefebvre, J. Yuan and R. L. Johnston, *Sci. Rep.*, 2016, **6**, 35226.
- 35 I. Demiroglu, Z. Y. Li, L. Piccolo and R. L. Johnston, *Catal. Sci. Technol.*, 2016, **6**, 6916–6931.
- 36 I. Demiroglu, Z. Y. Li, L. Piccolo and R. L. Johnston, *Comput. Theor. Chem.*, 2017, **1107**, 142–151.
- 37 H. Li and P. Afanasiev, *Mater. Res. Bull.*, 2011, **46**, 2506–2514.
- 38 A. De Clercq, W. Dachraoui, O. Margeat, K. Pelzer, C. R. Henry and S. Giorgio, *J. Phys. Chem. Lett.*, 2014, **5**, 2126–2130.
- 39 B. J. Berne and R. Pecora, *Dynamic Light Scattering: With Applications to Chemistry, Biology and Physics*, Dover Publications, 2000.
- 40 B. Wiley, T. Herricks, Y. Sun and Y. Xia, *Nano Lett.*, 2004, **4**, 1733–1739.
- 41 J. M. Yuk, J. Park, P. Ercius, K. Kim, D. J. Hellebusch, M. F. Crommie, J. Y. Lee, A. Zettl and A. P. Alivisatos, *Science*, 2012, **336**, 61–64.
- 42 M. Kahlweit, *Adv. Colloid Interface Sci.*, 1975, **5**, 1–35.
- 43 E. Csapó, A. Oszkó, E. Varga, Á. Juhász, N. Buzás, L. Kőrösi, A. Majzik and I. Dékány, *Colloids Surf., A*, 2012, **415**, 281–287.
- 44 M. Pradhan, S. Sarkar, A. K. Sinha, M. Basu and T. Pal, *J. Phys. Chem. C*, 2010, **114**, 16129–16142.
- 45 N. Toshima, *Pure Appl. Chem.*, 2000, **72**, 317–325.
- 46 Y. Zhao, L. Jia, J. A. Medrano, J. R. H. Ross and L. Lefferts, *ACS Catal.*, 2013, **3**, 2341–2352.
- 47 K. Hadjiivanov, *Appl. Surf. Sci.*, 1998, **135**, 331–338.
- 48 F. Boccuzzi and A. Chiorino, *J. Phys. Chem. B*, 2000, **104**, 5414–5416.
- 49 H. Klimev, K. Fajerweg, K. Chakarova, L. Delannoy, C. Louis and K. Hadjiivanov, *J. Mater. Sci.*, 2007, **42**, 3299–3306.
- 50 L. Piccolo, H. Daly, A. Valcarcel and F. C. Meunier, *Appl. Catal., B*, 2009, **86**, 190–195.
- 51 M. Primet, *J. Chem. Soc., Faraday Trans. 1*, 1978, **74**, 2570–2580.
- 52 J. T. Yates, T. M. Duncan, S. D. Worley and R. W. Vaughan, *J. Chem. Phys.*, 1979, **70**, 1219–1224.
- 53 F. Solymosi and M. Pasztor, *J. Phys. Chem.*, 1985, **89**, 4789–4793.
- 54 L. L. Sheu and W. M. H. Sachtler, *J. Mol. Catal.*, 1993, **81**, 267–278.
- 55 S. Trautmann and M. Baerns, *J. Catal.*, 1994, **150**, 335–344.
- 56 C. Langlois, Z. L. Li, J. Yuan, D. Alloyeau, J. Nelayah, D. Bochicchio, R. Ferrando and C. Ricolleau, *Nanoscale*, 2012, **4**, 3381–3388.
- 57 J. Šimonik and L. Beránek, *J. Catal.*, 1972, **24**, 348–351.
- 58 H. Noller and W. M. Lin, *J. Catal.*, 1984, **85**, 25–30.
- 59 J. M. Campelo, A. Garcia, D. Luna and J. M. Marinás, *J. Catal.*, 1988, **113**, 172–184.
- 60 M. A. Vannice and B. Sen, *J. Catal.*, 1989, **115**, 65–78.
- 61 C. Milone, M. C. Trapani and S. Galvagno, *Appl. Catal., A*, 2008, **337**, 163–167.
- 62 J. Shi, R. Nie, P. Chen and Z. Hou, *Catal. Commun.*, 2013, **41**, 101–105.
- 63 P. N. Rylander and N. Himmelstein, *Engelhard Ind., Tech. Bull.*, 1964, **4**, 131–134.